

ANSWERS TO LEWIS ACIDS AND BASES

- (a) base (b) base (c) acid (d) acid (e) acid (f) acid (g) acid (h) base
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| (a) acid = CO_2 , base = O^{2-} | (d) acid = H^+ , base = F^- |
| (b) acid = Cu^{2+} , base = H_2O | (e) acid = AlH_3 , base = NH_3 |
| (c) acid = Ag^+ , base = CN^- | (f) acid = CO_2 , base = CN^- |
- One Br on each AlBr_3 molecule donates a pair of electrons into a vacant Al orbital on the other molecule.